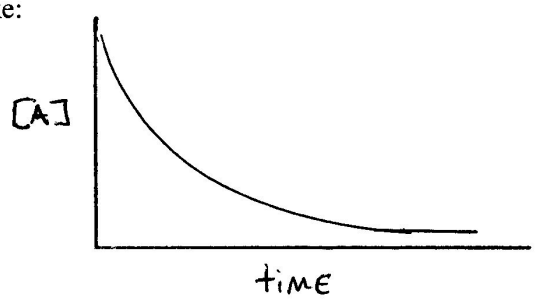
1cont.) For each statement, indicate if it is describing a first order, second order, or zero order reaction.

E) The units for k for this order could be mol/L-sec.

F) A plot of 1/[A] vs time gives a straight line with a positive slope.

G) For this order, the half-life gets shorter as the initial concentration increases.

H) For this order, a plot of [A] vs time would look like:



I) For this order, the rate and the rate constant stay the same throughout the course of the reaction

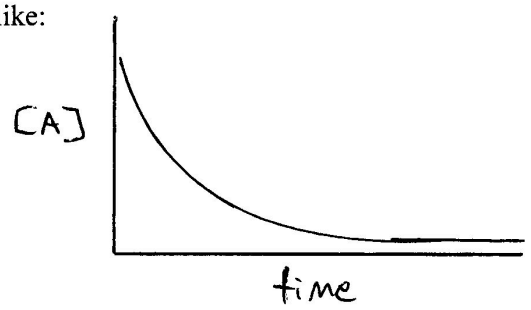
1cont.) For each statement, indicate if it is describing a first order, second order, or zero order reaction.

E) The units for k for this order could be mol/L-sec.

F) A plot of 1/[A] vs time gives a straight line with a positive slope.

G) For this order, the half-life gets shorter as the initial concentration increases.

H) For this order, a plot of [A] vs time would look like:



I) For this order, the rate and the rate constant stay the same throughout the course of the reaction