

Please complete the following double replacement reactions.
You should be attempting to write all equations as net ionic equations.

1. A solution of acetic acid is poured over a pile of potassium hydrogen carbonate.
2. Aqueous solutions of copper(II) chloride and potassium sulfide are combined.
3. A solution of rubidium hydroxide is poured over a pile of solid ammonium oxalate.
4. A solution of hydroiodic acid is combined with nickel(II)sulfite powder.
5. Solutions of molybdenum(IV)perchlorate and osmium(II)acetate are combined.
6. Ammonia gas is bubbled through a solution of hydrobromic acid.
7. Solutions of ammonium phosphate and sodium hydroxide are combined.
8. A solution of hydrofluoric acid is added to solid strontium carbonate.
9. Hypoiodic acid is combined with solid magnesium hydroxide.
10. Magnesium carbonate powder is added to a solution of oxalic acid.
11. A solution of hydrobromic acid is combined with solid silver sulfite.
12. Solutions of sulfuric acid and barium hydroxide are combined.
13. Carbon dioxide gas is bubbled through a solution of cesium hydroxide.
14. Carbon dioxide gas is bubbled into a solution of titanium(II)chloride
15. Lithium hydroxide powder as added to a solution of sulfurous acid.
16. Solutions of calcium chloride and ammonium sulfide are combined.
17. A solution of acetic acid is poured over solid strontium hydroxide pellets.
18. Ammonia gas is bubbled through a solution of bromic acid.
19. Solutions of tin(IV)chloride and rubidium dichromate are combined.
20. Solutions of iron(III)nitrate and lithium hydroxide are combined.
21. A solution containing two moles of phosphoric acid is combined with a solution containing 3 moles of calcium hydroxide.
22. A solution of nitric acid is combined with a solution of sodium fluoride.
23. Sulfur dioxide gas is bubbled through a solution of potassium hydroxide.
24. A calcium hydroxide solution as added to a pile of ammonium dichromate.
25. Carbon dioxide gas is bubbled through a solution of sodium carbonate.

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