

Label the following formulae according to the following codes:

- **I = ionic . . . [BI = binary ionic . . . TI = ternary ionic (polyatomic ions)]**
- **M = molecular . . .**
- **A = acid . . . [BA = binary acid (or non-oxyacid) . . . OA = oxyacid]**

Part I: Give the correct formula for each of the following . . .

1. mercury(II)sulfate_____
2. lithium acetate_____
3. zinc oxide_____
4. sulfuric acid_____
5. diphosphorus pentoxide_____
6. barium phosphate_____
7. ammonium hydroxide_____
8. iron(II)perchlorate_____
9. potassium fluoride_____
10. perchloric acid_____
11. sodium carbonate_____
12. nickel(II)oxide_____
13. boron trifluoride_____
14. magnesium phosphate_____
15. hydrogen chloride_____
16. hydrobromic acid_____
17. calcium bromide_____
18. carbon dioxide_____
19. lead(IV)nitrate_____
20. sodium hydroxide_____
21. aluminum hydrogen sulfate_____
22. hydroiodic acid_____
23. nitrous acid_____
24. copper(II)nitride_____
25. carbon tetrabromide_____

Part II: Give the correct name for each of the following . . .

26. Al_2O_3 _____
27. Ag_2S _____
28. CBr_4 _____
29. PbSO_3 _____
30. LiI _____
31. $\text{Mg}(\text{OH})_2$ _____
32. K_2S _____
33. $\text{Ca}(\text{CN})_2$ _____
34. $\text{Fe}_3(\text{PO}_4)_2$ _____
35. N_2O_5 _____
36. HBrO_2 _____
37. HCl (aq)_____
38. HF (aq)_____
39. KMnO_4 _____
40. OF_2 _____
41. CO_2 _____
42. $\text{Al}(\text{NO}_2)_3$ _____
43. CuO _____
44. Be_3P_2 _____
45. NaF _____
46. CaHPO_4 _____
47. Al_2S_3 _____
48. Cu_2O _____
49. Br_2O_7 _____
50. SnCl_4 _____